

Homework 3 Solutions

2. bcc: $U(R) = 2N\epsilon[9.114(\sigma/R)^{12} - 12.253(\sigma/R)^6]$. At equilibrium $R_0^6 = 1.488\sigma^6$, and $U(R_0) = 2N\epsilon(-2.816)$.

fcc: $U(R) = 2N\epsilon[12.132(\sigma/R)^{12} - 14.454(\sigma/R)^6]$. At equilibrium $R_0^6 = 1.679\sigma^6$, and $U(R_0) = 2N\epsilon(-4.305)$. Thus the cohesive energy ratio $\text{bcc/fcc} = 0.956$, so that the fcc structure is more stable than the bcc.

$$\begin{aligned} 3. \quad |U| &= 8.60 N\epsilon \\ &= (8.60)(6.02 \times 10^{23})(50 \times 10^{-16}) = 25.9 \times 10^9 \text{ erg/mol} \\ &= 2.59 \text{ kJ/mol.} \end{aligned}$$

This will be decreased significantly by quantum corrections, so that it is quite reasonable to find the same melting points for H_2 and Ne.

5a.

$$U(R) = N \left(\frac{A}{R^n} - \frac{\alpha q^2}{R} \right); \quad \alpha = 2 \log 2 = \text{Madelung const.}$$

In equilibrium

$$\frac{\partial U}{\partial R} = N \left(-\frac{nA}{R_0^{n+1}} + \frac{\alpha q^2}{R_0^2} \right) = 0; \quad R_0^n = \frac{nA}{\alpha q^2},$$

and

$$U(R_0) = -\frac{N\alpha q^2}{R_0} \left(1 - \frac{1}{n}\right).$$

$$b. \quad U(R_0 - R_0\delta) = U(R_0) + \frac{1}{2} \frac{\partial^2 U}{\partial R^2} R_0 (R_0\delta)^2 + \dots,$$

bearing in mind that in equilibrium $(\partial U / \partial R)_{R_0} = 0$.

$$\left(\frac{\partial^2 U}{\partial R^2} \right)_{R_0} = N \left(\frac{n(n+1)A}{R_0^{n+2}} - \frac{2\alpha q^2}{R_0^3} \right) = N \left(\frac{(n+1)\alpha q^2}{R_0^3} - \frac{2\alpha q^2}{R_0^3} \right)$$

For a unit length $2NR_0 = 1$, whence

$$\left(\frac{\partial^2 U}{\partial R^2} \right)_{R_0} = \frac{\alpha q^2}{2R_0^4} (n-1); \quad C = R_0^2 \left. \frac{\partial^2 U}{\partial R^2} \right|_{R_0} = \frac{(n-1)q^2 \log 2}{R_0^2}.$$

8. From (37) we have $e_{xx} = S_{11}X_x$, because all other stress components are zero. By (51), $3S_{11} = 2/(C_{11} - C_{12}) + 1/(C_{11} + C_{12})$.

Thus $Y = (C_{11}^2 + C_{12}C_{11} - 2C_{12}^2)/(C_{11} + C_{12})$;

further, also from (37), $e_{yy} = S_{21}X_x$,

whence $\sigma = e_{yy}/e_{xx} = S_{21}/S_{11} = -C_{12}/(C_{11} + C_{12})$.